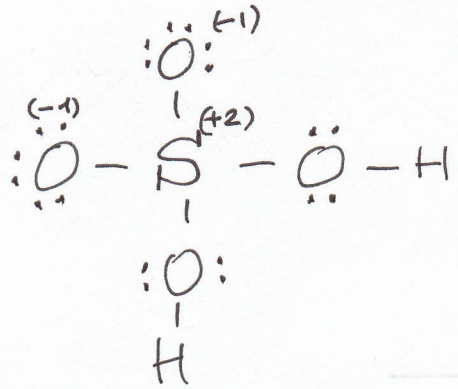
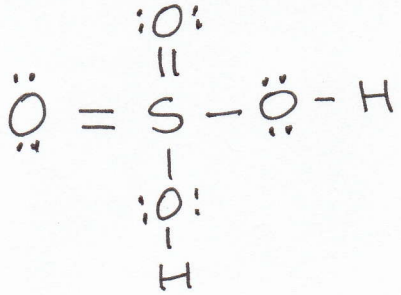
Acido solforico H_2SO_4

H	2	· 1	=	2
S	1	· 6	=	6
O	4	· 6	=	24
				32 $\angle 2$ = 16 doppietti





Per la VSEPR qui dobbiamo considerare:

a) lo zolfo AB_4

b) i due atomi di ossigeno degli ossidrili

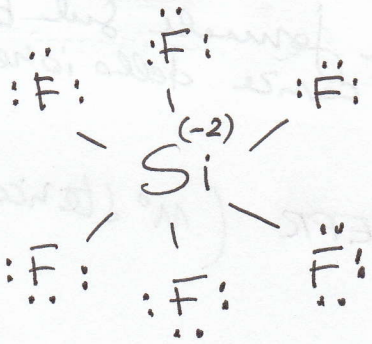
Per quanto riguarda S la geometria è del tipo AB_4 quindi è tetraedrica

Per quanto riguarda l'O di OH esso ha 2 doppietti e 2 sostituenti, quindi è del tipo AB_2E_2, cioè angolare (come in H_2O)

Ione esafluorosilicico SiF_6^{2-}

Si	1	4	=	4
F	6	7	=	42
2e				2

$\frac{48}{18} L2 = \cancel{24}$

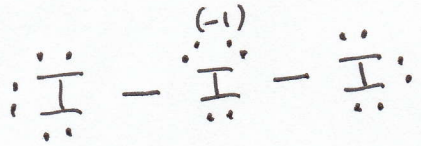


AB_6 (ottaedrica)

Ione triioduro (I_3^-)

I	3	· 7	=	21
e⁻				1

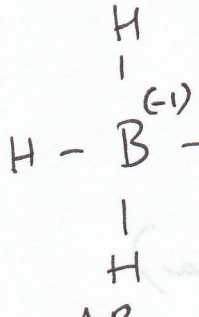
$\frac{22}{11} L2$ doppietti



AB_2E_3 lineare

Ione tetraidroborato BH_4^-

B	1	3	=	3
H	4	1		4
e^-				1
				8 | 2 = 4



AB_4

strutture ok! xché il boro ha raggiunto l'ottetto

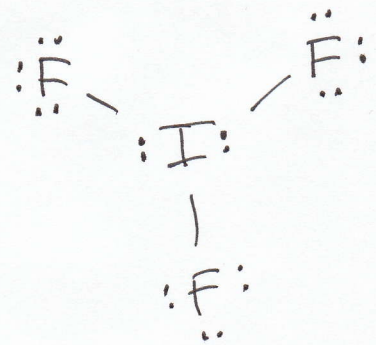
Struttura tetraedrica secondo la VSEPR (N° sterico = 4)

---

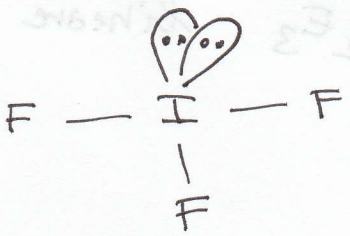
o

Trifluoruro di iodio IF_3

I	1	· 7	=	7
F	3	· 7	=	21
				28 | 2 = 14

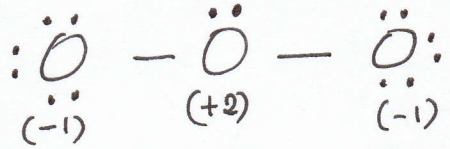


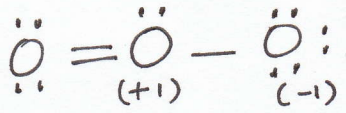
AB_3E_2



Ozono O_3

O 3 · 6 = 18 $\not{2}$ = 9 doppietti



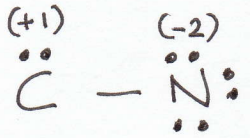


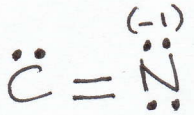
ok! xché l'atomo centrale ha raggiunto l'ottetto (l'ossigeno fa parte del II periodo!)

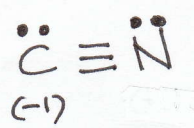
AB_2E angolare (a V)

Ione cianuro CN^-

C	1 · 4 =	4
N	1 · 5 =	5
1e⁻		1
		10 $\not{2}$ = 5



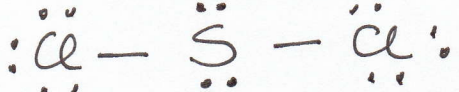




ok! il carbonio ha l'ottetto completo!

SCl_2 dicloruro di zolfo

S	$1 \cdot 6$	$= 6$
Cl	$2 \cdot 7$	$= 14$
		$20 \quad 2 = 10$



AB_2E_2 angolare (a V)

NH_2OH idrossilammina

N	$1 \cdot 5$	=	5
H	$3 \cdot 1$	=	3
O	$1 \cdot 6$	=	6
			14 L2 = 7